



Assessment = ★



**Apply:**  
SC17 Groups in the  
periodic table

## Electronic Configuration

- 1<sup>st</sup> shell- max 2 electrons
- 2<sup>nd</sup> shell- max 8 electrons
- 3<sup>rd</sup> shell- max 8 electrons
- 4<sup>th</sup> shell- max 18 electrons



TIP: Always put  
your electrons  
on in the order  
North, East,  
South, West to  
avoid later  
confusion!



## Periodic Table of the Elements

## Assessment & Review



## Revision lesson

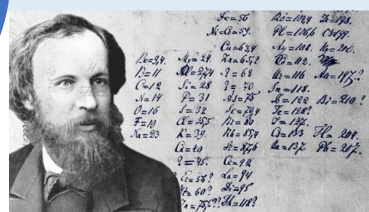
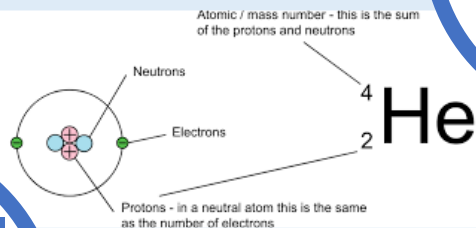
Focus on retrieval,  
key word  
definitions and  
equations.

## Electronic configurations

Electron shells,  
configurations and  
the link to the  
periodic table

## Atomic number & the periodic table

The modern periodic  
table compared to  
Mendeleev's table

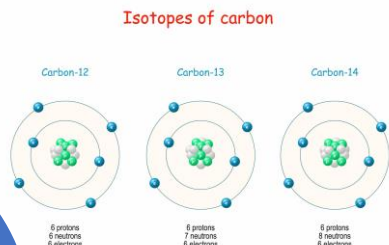


## Elements & the periodic table

Mendeleev's table

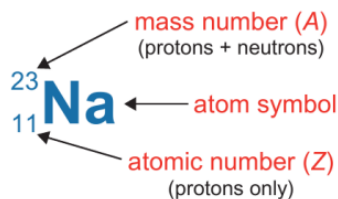
## Isotopes

Examples of  
isotopes, **Higher** –  
calculating relative  
atomic mass



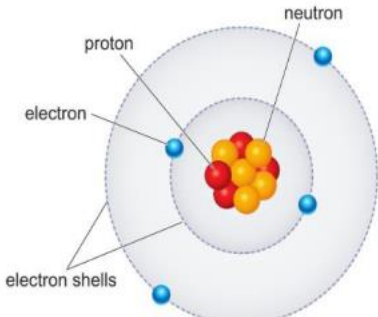
## Atomic number, mass number

Mass mainly in  
nucleus, calculating  
Protons/neutrons/  
electrons



## Structure of an atom

Model of the atom,  
sub-atomic  
particles



**Retrieve:**  
C1 1.1 The particle  
model  
C2 1 The periodic  
table  
SC2 States of matter

## Keywords

Atoms, element, subatomic particle, proton, neutron, electron, nucleus, electron shell, atomic number, mass number, isotopes, relative atomic mass, periodic table, properties, period, group, electronic configuration.